

Tittle: Crystallisation of oxalic acid

Work instructions

Task: Crystallise oxalic acid from its aqueous solution.

Theory

Oxalic acid (HOOC-COOH) is the simplest dicarboxylic acid. It is a white, relatively watersoluble substance. It forms nicely developed rectangular crystals. They can be obtained by very slow cooling of a supersaturated solution.

The advantage of this experiment is its low consumption of chemical, since the crystallised oxalic acid can be reused, also for further experiments.

Equipment: beaker 500 ml, laboratory spoon, glass rod, electric cooker, distilled water

Chemicals: oxalic acid dihydrate

Procedures:

- 1. Pour 250 ml of distilled water into a 500 ml beaker.
- 2. To this beaker add about 31.5 g of oxalic acid dihydrate with a spoon. Wait for some of the solid chemical to dissolve and then slowly heat the solution.
- 3. When all of the oxalic acid has dissolved, add the 31.5 g of solid oxalic acid dihydrate again. Then wait until this proportion has also dissolved.
- 4. Stop heating and place the beaker in a safe place, preferably inside a laboratory bench.
- 5. Allow to crystallise freely for at least three days.
- 6. When this process is complete, observe the transparent crystals that have formed. Most of them are rectangular in shape and you can pick them out with tweezers on a colored pad to see their shape.

Chemicals	Form	H-statements	P-statements
$C_2H_2O_4$ ·2H_2O	Solid	H302 + H312, H318	P280, P305 + P351 + P338

Management of chemical substances

Name of the	the Digitization of chemistry experiments to improve the quality	
project:	support chemistry teaching in secondary schools	
Acronym:	ChemIQSoc	
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number:		



Sources of risk and assessment of risk severity

Oxalic acid is harmful if swallowed or in contact with skin, eye damage may occur. There is a possibility of scalding and sunburn.

Waste management method

Certified chemical waste disposal company.

Risk reduction measures

Lab coat, goggles, gloves.